

Note: the following are examples of the type of calculations you may be asked to perform in the TE4 narrative paper. This is not a complete exam paper. Calculation questions typically make up 20 out of 40 marks for a paper.

EXAMPLE CALCULATIONS ASSOCIATED WITH THE MEASUREMENT OF GASES AND VAPOURS USING INSTRUMENTAL TECHNIQUES

1 To calculate total NO_x in mg/m³ at reference conditions given NO and NO₂ measurements in ppm

NO concentration = 34 ppm (dry)
NO₂ concentration = 15 ppm (dry)
Measured oxygen level = 12%
Reference conditions = STP, 15% O₂, dry
Atomic weight of N = 14
Atomic weight of O = 16
Molar volume = 22.4 litres

1.1 To calculate the total NO_x concentration in ppm

Total NO_x concentration = NO concentration + NO₂ concentration

= 34 ppm + 15 ppm

= 49 ppm

1.2 To calculate the molecular weight of NO_x as NO₂

Molecular weight of NO₂ = atomic weight of N + molecular weight of O₂

= 14 + (16 x 2)

= 46

1.3 To convert NO_x concentrations in ppm to NO₂ in mg/m³

NO_x concentration = 49 ppm
Molecular weight of NO₂ = 46
Molar volume = 22.4 litres

Concentration in mg/m³ = $\frac{\text{concentration (ppm)} \times \text{molecular weight of substance}}{\text{molar volume}}$

= $\frac{49 \text{ ppm} \times 46}{22.4}$

= 100.6 mg/m³

1.4 To calculate the concentration of NO₂ at reference conditions

Measured oxygen level = 12%

Reference oxygen = 15%

$$\begin{aligned}\text{Oxygen correction factor} &= \frac{(21 - \text{reference oxygen})}{(21 - \text{measured oxygen})} \\ &= \frac{(21 - 15)}{(21 - 12)} \\ &= 0.67\end{aligned}$$

Concentration at reference conditions

$$\begin{aligned}&= \text{Concentration as measured} \times \text{correction factor for oxygen} \\ &= 100.6 \text{ mg/m}^3 \times 0.67 \\ &= 67.4 \text{ mg/m}^3\end{aligned}$$

2 To calculate concentrations at reference conditions in mg/m³ from analyser results in ppm

Concentration of substance = 120 ppm

Measured moisture level = 9%

Reference moisture is dry

Measured oxygen level = 12.5% (wet)

Reference oxygen = 11%

Molecular weight of substance = 28

Molar volume = 22.4 litres

(In this example the molecule is CO)

2.1 To convert concentrations in ppm to mg/m³

Concentration of substance = 120 ppm

Molecular weight of substance = 28

Molar volume = 22.4 litres

$$\begin{aligned}\text{Concentration in mg/m}^3 &= \frac{\text{concentration (ppm)} \times \text{molecular weight of substance}}{\text{molar volume}} \\ &= \frac{120 \text{ ppm} \times 28}{22.4} \\ &= 150 \text{ mg/m}^3\end{aligned}$$

2.2 To calculate the concentration of a substance at reference conditions

$$\begin{aligned}\text{Moisture correction factor} &= \frac{100}{(100 - \text{measured moisture})} \\ &= \frac{100}{(100 - 9)} \\ &= 1.1\end{aligned}$$

(Note that the oxygen was measured on a wet basis and should be corrected to dry conditions)

$$\begin{aligned}\text{Oxygen at reference conditions} &= \text{measured oxygen} \times \text{correction factor for moisture} \\ &= 12.5\% \times 1.1 \\ &= 13.8\% \text{ (dry)}\end{aligned}$$

$$\begin{aligned}\text{Oxygen correction factor} &= \frac{(21 - \text{reference oxygen})}{(21 - \text{measured oxygen, dry})} \\ &= \frac{(21 - 11)}{(21 - 13.8)} \\ &= 1.4\end{aligned}$$

Concentration of substance at reference conditions

$$\begin{aligned}&= \text{concentration as measured} \times \text{correction factor for oxygen} \times \\ &\quad \text{correction factor for moisture} \\ &= 150 \text{ mg/m}^3 \times 1.4 \times 1.1 \\ &= 231 \text{ mg/m}^3\end{aligned}$$

3 To express analyser readings in ppm (wet) at reference conditions in mg/m³ (dry)

NO_x concentration = 80 ppm (wet)
Measured moisture level = 10%
Molar volume at STP = 22.4 litres
Atomic weight of N = 14
Atomic weight of O = 16

3.1 To calculate the molecular weight of NO₂

Molecular weight of NO₂ = atomic weight of N + molecular weight of O₂

= 14 + (16 x 2)

= 46

3.2 To convert concentration in ppm to mg/m³

Concentration in mg/m³ = $\frac{\text{concentration (ppm)} \times \text{molecular weight of substance}}{\text{molar volume}}$

= $\frac{80 \text{ ppm} \times 46}{22.4}$

= 164 mg/Nm³ (wet)

3.3 To calculate the concentration at reference conditions dry

Moisture correction factor = $\frac{100}{(100 - \text{measured moisture})}$

= $\frac{100}{100 - 10}$

= 1.1

NO₂ concentration (dry)

= concentration as measured x correction factor for moisture

= 164 mg/Nm³ x 1.1

= 180 mg/Nm³ (dry)

4 To calculate a mass emission rate in kg/h (Example 1)

Volume flow rate at STP (wet) = 43 Nm³/s

NO₂ concentration (wet) = 164 mg/m³

Seconds in 1 hour = 3600

4.1 To calculate the volume flow in the stack

Volume flow rate in Nm³/h = volume flow rate in Nm³/s x seconds in 1 hour

$$= 43 \text{ Nm}^3/\text{s} \times 3600$$

$$= 154,800 \text{ Nm}^3/\text{h}$$

4.2 To calculate the mass emission to atmosphere in kg/h

Mass emission rate in kg/h = $\frac{\text{volume flow rate m}^3/\text{h} \times \text{concentration in mg/m}^3}{1 \times 10^6}$

$$= \frac{154,800 \text{ Nm}^3/\text{h} \times 164 \text{ mg/m}^3}{1 \times 10^6}$$

$$= 25.4 \text{ kg/h}$$

(Note dividing by 1 x 10⁶ converts the result from mg to kg)

5 To calculate a mass emission in kg/h (Example 2)

Average stack gas velocity = 10 m/s

Stack diameter = 0.9 m (radius = 0.45 m)

Concentration of substance = 150 mg/m³ (measured wet with no correction for oxygen or temperature)

5.1 To calculate the cross section area of the stack

Cross sectional area of stack = πr^2

$$= 3.14 \times 0.45^2$$

$$= 0.64 \text{ m}^2$$

5.2 To calculate the volume flow in the stack

Volume flow rate in m³/s = cross section area of stack x average gas velocity

$$= 0.64 \text{ m}^2 \times 10 \text{ m/s}$$

$$= 6.4 \text{ m}^3/\text{s}$$

Volume flow rate in m³/h = volume flow rate in m³/s x seconds in 1 hour

$$= 6.4 \text{ m}^3/\text{s} \times 3600$$

$$= 23,040 \text{ m}^3/\text{h}$$

5.3 To calculate the mass emission in kg/h at stack conditions

$$\begin{aligned}\text{Mass emission} &= \frac{\text{volume flow rate} \times \text{concentration of substance}}{1 \times 10^6} \\ &= \frac{23,040 \text{ m}^3/\text{h} \times 150 \text{ mg}/\text{m}^3}{1 \times 10^6} \\ &= 3.46 \text{ kg}/\text{h}\end{aligned}$$

(Note dividing by 1×10^6 converts the result from mg to kg)

6 To convert VOC results from a FID analyser to reference conditions

VOC concentration = 23 mgC/m³

Measured moisture level = 9.5%

Reference moisture is dry

Measured oxygen level = 13.2% (dry)

Reference oxygen = 11%

(Note that the FID analyser measures hot and wet)

$$\begin{aligned}\text{Oxygen correction factor} &= \frac{(21 - \text{reference oxygen})}{(21 - \text{measured oxygen})} \\ &= \frac{(21 - 11)}{(21 - 13.2)} \\ &= 1.3\end{aligned}$$

$$\begin{aligned}\text{Moisture correction factor} &= \frac{100}{(100 - \text{measured moisture})} \\ &= \frac{100}{(100 - 9.5)} \\ &= 1.1\end{aligned}$$

Concentration at reference conditions

$$\begin{aligned}&= \text{Concentration as measured} \times \text{correction factor for oxygen} \times \text{correction factor for moisture} \\ &= 23 \text{ mgC}/\text{m}^3 \times 1.3 \times 1.1 \\ &= 33 \text{ mgC}/\text{m}^3\end{aligned}$$

7 To calculate the concentration of VOCs as Carbon (C) as a dry gas and calculate the mass emissions of carbon and toluene from the results provided by a FID analyser

VOC concentration = 185 ppm (propane equivalent C_3H_8)

Measured moisture level = 6.5%

Atomic weight of C = 12

Atomic weight of H = 1

Molar volume at STP = 22.4 litres

Stack gas flow rate = 0.4 m³/s

7.1 To calculate the molecular weight of propane as carbon

Molecular weight of propane as carbon

= number of carbon atoms in molecule x atomic weight of carbon

= 3 x 12

= 36

7.2 To calculate the concentration of VOC in mgC at STP wet

= $\frac{\text{concentration in ppm} \times \text{molecular weight of propane as carbon}}{\text{molar volume}}$

= $\frac{185 \text{ ppm} \times 36}{22.4}$

= 297 mgC/Nm³

7.3 To calculate the concentration of VOC in mgC as a dry gas

Moisture correction factor = $\frac{100}{(100 - \text{measured moisture})}$

= $\frac{100}{100 - 6.5}$

= 1.07

Concentration of VOC(dry) = VOC concentration wet x correction factor for moisture

= 297 mgC/Nm³ x 1.07

= 318 mgC/Nm³

7.4 To calculate the hourly emission rate in gC/h

Stack gas flow rate = 0.4 m³/s

Concentration of VOC = 297 mgC/Nm³ (wet)

$$\begin{aligned}\text{Volume flow rate in m}^3/\text{h} &= \text{flow rate in m}^3/\text{s} \times \text{seconds in 1 hour} \\ &= 0.4 \text{ m}^3/\text{s} \times 3600 \\ &= 1440 \text{ m}^3/\text{h}\end{aligned}$$

$$\begin{aligned}\text{Mass emission rate g/h} &= \frac{\text{volume flow rate m}^3/\text{h} \times \text{concentration of substance}}{1000} \\ &= \frac{1440 \text{ m}^3/\text{h} \times 297 \text{ mgC/Nm}^3}{1000} \\ &= 428 \text{ gC/h}\end{aligned}$$

(Note dividing by 1000 converts the result from mg to g)

7.5 To calculate the mass emission as g toluene per hour

Toluene = C₇H₈

Molecular weight of toluene as mgC

$$\begin{aligned}&= \text{number of carbon atoms in molecule} \times \text{atomic weight of carbon} \\ &= 7 \times 12 \\ &= 84\end{aligned}$$

Molecular weight of toluene

$$\begin{aligned}&= (\text{number of carbon atoms} \times \text{molecular weight of carbon}) + \\ &\quad (\text{number of hydrogen atoms} \times \text{atomic weight of hydrogen}) \\ &= (7 \times 12) + (8 \times 1) \\ &= 92\end{aligned}$$

Ratio of molecular weight of toluene to molecular weight of toluene as carbon

$$\begin{aligned}&= \frac{\text{molecular weight of toluene}}{\text{molecular weight of toluene as carbon}} \\ &= \frac{92}{84} \\ &= 1.095\end{aligned}$$

Mass emission expressed as gT/h

$$= \text{mass emission rate gC/h} \times \text{toluene to carbon mass ratio}$$

$$= 428 \text{ gC/h} \times 1.095$$

$$= 469 \text{ gT/h}$$

8 To calculate an uncertainty budget

The performance characteristics of the analyser are as follows:

Type

B Sensitivity to ambient temperature = 1.5% of the certified range per 10K of variation

B Sensitivity to atmospheric pressure = 0.7% of the certified range per kPa

B Uncertainty of calibration gas = 1.5%

A Standard deviation of repeatability at span = 0.9% of the certified range

Certification range of analyser = 0 – 175 mg/m³

Maximum ambient temperature variation during sampling = 8K (± 4K relative to the temperature value at calibration)*

Ambient pressure range during sampling = 98.8 – 101.2 kPa (± 1.2 kPa relative to a pressure value of 100 kPa at calibration)*

Calibration gas concentration = 140 mg/m³

* Note : The exams will not require a calculation of an uncertainty where the variation is not symmetrical about the value at calibration

8.1 To calculate the standard uncertainty in mg/m³ associated with the ambient temperature variations

Sensitivity of analyser to ambient temperature variations = 1.5% of the certified range per 10K of variation

Certification range = 0 – 175 mg/m³

Maximum ambient temperature variation during sampling = ± 4 K

8.2 To calculate the sensitivity per degree of ambient temperature change

$$= \frac{\text{certified range} \times \text{sensitivity}}{10 \text{ K}}$$

$$= \frac{175 \text{ mg/m}^3 \times 0.015}{10\text{K}}$$

$$= 0.26 \text{ mg/m}^3 \text{ per degree}$$

8.3 To calculate the standard uncertainty associated with the ambient temperature variation

Standard uncertainty U_{temp}

$$\begin{aligned} &= \frac{\text{sensitivity per degree} \times \text{variation in ambient temperature}}{\sqrt{3}} \\ &= \frac{0.26 \text{ mg/m}^3 \times 4 \text{ K}}{\sqrt{3}} \\ &= 0.6 \text{ mg/m}^3 \end{aligned}$$

8.4 To calculate the standard uncertainty in mg/m³ associated with ambient pressure changes

Sensitivity of analyser to ambient pressure variations = 0.7% of the certified range per kPa of pressure change

Certification range = 175 mg/m³

Ambient pressure range during sampling = 98.8 – 101.2 kPa

$$\begin{aligned} \text{Variation in ambient pressure} &= \frac{101.2 - 98.8}{2} \\ &= \pm 1.2 \text{ kPa} \end{aligned}$$

$$\begin{aligned} \text{Sensitivity per kPa of pressure change} &= \text{certification range} \times \text{sensitivity} \\ &= 175 \text{ mg/m}^3 \times 0.007 \\ &= 1.23 \text{ mg/m}^3 \text{ per kPa} \end{aligned}$$

Standard uncertainty associated with ambient pressure changes U_{apress}

$$\begin{aligned} &= \frac{\text{sensitivity} \times \text{variation in ambient pressure}}{\sqrt{3}} \\ &= \frac{1.23 \text{ mg/m}^3 \times 1.2 \text{ kPa}}{\sqrt{3}} \\ &= 0.85 \text{ mg/m}^3 \end{aligned}$$

8.5 To calculate the standard uncertainty of the calibration gas in mg/m³

Uncertainty of calibration gas = 1.5%

Calibration gas value = 140 mg/m³

Standard uncertainty of calibration gas U_{cal}

$$= \frac{\text{calibrated gas value} \times \text{uncertainty of calibration gas}}{\sqrt{3}}$$

$$= \frac{140 \text{ mg/m}^3 \times 0.015}{\sqrt{3}}$$
$$= 1.21 \text{ mg/m}^3$$

8.6 To calculate the standard deviation of repeatability at span level in mg/m³

Standard deviation of repeatability at span level = 0.9% of the certified range (note that this is a 'Type A' uncertainty and does not need to be divided by $\sqrt{3}$)

Certified range of analyser = 175 mg/m³

Standard deviation of repeatability at span $U_{s,r}$

$$= \text{certified range} \times \text{standard deviation of repeatability at span}$$
$$= 175 \text{ mg/m}^3 \times 0.9\%$$
$$= 1.58 \text{ mg/m}^3$$

8.7 To calculate the combined standard uncertainty

$$\text{Combined uncertainty} = \sqrt{U_{\text{temp}}^2 + U_{\text{apress}}^2 + U_{\text{cal}}^2 + U_{s,r}^2}$$
$$= \sqrt{0.6^2 + 0.85^2 + 1.21^2 + 1.58^2}$$
$$= \sqrt{0.36 + 0.72 + 1.46 + 2.5}$$
$$= 2.25 \text{ mg/m}^3$$

$$\text{Expanded uncertainty} = \text{combined uncertainty} \times 1.96$$
$$= 2.25 \text{ mg/m}^3 \times 1.96$$
$$= 4.4 \text{ mg/m}^3$$